

Charge as an ion.

lose electron. $+1$

gain electron $+2$

Groups 1-18 # of electrons on outer shell.

Periodic Table of the Elements

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1	2																
3	4																
5	6																
7	8																
9	10																
11	12																
13	14																
15	16																
17	18																

of Electron Shells.

Periods 1-7.

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Same number of shells, different number of protons.
 = More protons = ↑ electron affinity = more reactive easier to gain = More REACTIVE.
 Same shells = ↑ electron affinity = More REACTIVE.

Same number of shells, different number of protons = More protons & electron affinity (harder to lose) = less protons & electron affinity = more REACTIVE.
 More shells = ↓ electron affinity (easier to lose) = more REACTIVE.

 Metalloid
 Non-metal
 Metal