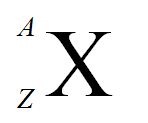
**Year 9 Chemistry Mid Topic Test Revision**

1. What do the A and Z represent in the following diagram?



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|  |

1. How is the number of neutrons in an atom calculated?

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|  |
|  |

1. State the charge of each of the following

|  |  |
| --- | --- |
| Protons |  |
| Neutrons |  |
| Electrons |  |

1. Complete the following table

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| Name of Element | Symbol | Number of protons | Number of neutrons | Number of Electrons (in neutral atom) |
| Helium |  |  |  |  |
| Aluminium |  |  |  |  |
| Hydrogen |  |  |  |  |
| Calcium |  |  |  |  |
| Lithium |  |  |  |  |

1. Complete the following table

|  |  |  |  |
| --- | --- | --- | --- |
| Symbol | Number of Electrons | Electron Configuration | Does this species have the same number of electrons as the neutral atom or has it gained or lost some |
| Cu |  |  |  |
| Cu2+ |  |  |  |
| H+ |  |  |  |
| Cl |  |  |  |

1. On the periodic table what is a group and a period?

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| --- |
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|  |

1. Explain how the groups of the periodic table relate to the valency of an ion?

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| --- |
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|  |
|  |

1. Fill out the following table

|  |  |  |
| --- | --- | --- |
| **Group Name** | **Group Number(s)** | **Properties** |
|  |  |  |
|  |  |  |
|  |  |  |
|  |  |  |
|  |  |  |

1. State the six rules for naming and writing ionic formula

|  |  |
| --- | --- |
|  | **Rule** |
| 1 |  |
| 2 |  |
| 3 |  |
| 4 |  |
| 5 |  |
| 6 |  |

**REVISE VALENCIES!!**