Table 25.4 The solubilities of some common salts

Al	$1 \text{ K}^{1+}$ , $Na^{1+}$ , $NH_4^{1+}$ salts are soluble.
Δl	NO; salts are soluble.
A	I Cl <sup>1-</sup> , Br <sup>1-</sup> , I <sup>1-</sup> salts are soluble, except Ag <sup>1+</sup> and Pb <sup>2+</sup> salts of these ions
Αl	$1 \text{ SO}_2^{2-}$ salts are soluble, except Ba <sup>2+</sup> , Pb <sup>2+</sup> , Ca <sup>2+</sup> and Ag <sup>1+</sup> salts of these
ion	ns.
Αl	$CO_3^{2-}$ salts are insoluble, except Na <sup>1+</sup> , K <sup>1+</sup> and NH <sub>4</sub> <sup>1+</sup> salts of this ion.
All	OH <sup>-</sup> compounds are insoluble, except Na <sup>1+</sup> , $K^{1+}$ , Ba <sup>2+</sup> and NH <sup>1+</sup>
	mpounds of this ion.
ΔII	$S^{2-}$ salts are insoluble, except Na <sup>1+</sup> , K <sup>1+</sup> and NH <sup>1+</sup> salts of this ion.
2	CaSO <sub>4</sub> and Ca(OH) <sub>2</sub> are slightly soluble.
Ь	PbCl $_{\rm z}$ and Pbl $_{\rm z}$ are siightly soluble.
C	NH <sub>a</sub> OH exists only as solution of NH <sub>a</sub> in water.

## Table 25.4 The solubilities of some common salts

A	ll K <sup>1+</sup> , Na <sup>1+</sup> , NH $_4^{1+}$ salts are soluble.		
Al	! NO <sup>1</sup> salts are soluble.		
Al	I CI <sup>1-</sup> , Br <sup>1-</sup> , I <sup>1-</sup> salts are soluble, except Ag <sup>1+</sup> and Pb <sup>2+</sup> salts of these ions		
Αl	$I SO_2^{2-}$ salts are soluble, except Ba <sup>2+</sup> , Pb <sup>2+</sup> , Ca <sup>2+</sup> and Ag <sup>1+</sup> salts of these		
io	ns.		
Al	$ICO_{\frac{1}{2}}^{2}$ salts are insoluble, except Na <sup>1+</sup> , K <sup>1+</sup> and NH <sub>4</sub> <sup>1+</sup> salts of this ion.		
Al	All OHT compounds are insoluble, except Na <sup>1+</sup> , K <sup>1+</sup> , Ba <sup>2+</sup> and NH <sup>1+</sup> <sub>2</sub>		
CO	compounds of this ion.		
Al	I S <sup>2+</sup> salts are insoluble, except Na <sup>1+</sup> , K <sup>1+</sup> and NH $_4^{1+}$ salts of this ion.		
a	$CaSO_4$ and $Ca(OH)_2$ are slightly soluble.		
b	$PbCl_2$ and $Pbl_2$ are slightly soluble.		
_	NH <sub>2</sub> OH exists only as solution of NH <sub>2</sub> in water.		

## Table 25.4 The solubilities of some common salts

1	All K1+, Na1+, NH4+ salts are soluble.	
2	All NO: salts are soluble.	
3	All C! $^{1}$ -, Br $^{1}$ -, $1^{1}$ - salts are soluble, except Ag $^{1}$ + and Pb $^{2}$ + salts of these ions.	
4	All $SO_4^{2-}$ salts are soluble, except $Ba^{2+}$ , $Pb^{2+}$ , $Ca^{2+}$ and $Ag^{1+}$ salts of these	
	ions.	
5	All $CO_3^{2-}$ salts are insoluble, except $Na^{1+}$ , $K^{1+}$ and $NH_4^{1+}$ salts of this ion.	
6	All OH <sup>-</sup> compounds are insoluble, except Na <sup>1+</sup> , K <sup>1+</sup> , Ba <sup>2+</sup> and NH $_4^{1+}$ compounds of this ion.	
7	All $S^{2-}$ salts are insoluble, except $Na^{1+}$ , $K^{1+}$ and $NH_4^{1-}$ salts of this ion.	
8	a CaSO <sub>4</sub> and Ca(OH) <sub>2</sub> are slightly soluble.	
	b PbCl <sub>2</sub> and Pbl <sub>2</sub> are slightly soluble.	

NH,OH exists only as solution of NH, in water.